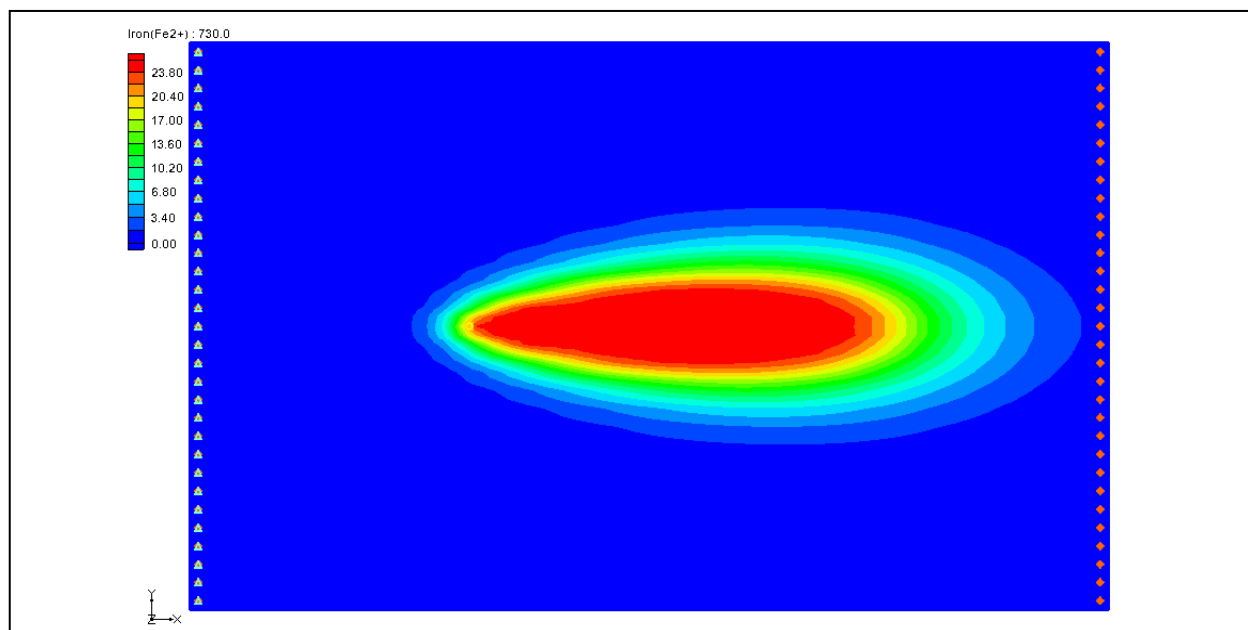


GMS 10.2 Tutorial

RT3D – BTEX Degradation with Multiple Electron Acceptors



Objectives

Use GMS and RT3D to model BTEX degradation using a multiple electron acceptor model.

Prerequisite Tutorials

- RT3D – Instantaneous Aerobic Degradation

Required Components

- Grid Module
- MODFLOW
- RT3D

Time

- 20–35 minutes



1	Introduction	2
1.1	Description of the Reaction Model.....	2
1.2	Description of Problem	5
1.3	Getting Started	6
2	Importing the MODFLOW Model	6
3	Building the Transport Model	6
4	Initializing the Model	7
5	The BTN Package	7
5.1	Starting Concentrations	7
5.2	Porosity	8
5.3	Stress Periods	8
5.4	Output Options	8
6	Assigning Concentrations to the Left Boundary	9
7	The Advection Package	9
8	The Dispersion Package	9
9	The Source/Sink Mixing Package	10
10	The Chemical Reaction Package	10
11	Running RT3D	11
12	Viewing the Results	12
12.1	Overlaying Contours	12
13	Other Post-Processing Options	14
14	Conclusion	14

1 Introduction

This tutorial illustrates the steps involved in using GMS and RT3D to model BTEX degradation using a multiple electron acceptor model. Since the flow model used in this simulation is similar to the flow model used in the “RT3D – Instantaneous Aerobic Degradation” tutorial, the steps involved in building the flow model will not be described in this tutorial. A predefined version of the flow model will be used.

This tutorial will begin with importing a MODFLOW model followed by defining RT3D inputs and boundary conditions, running RT3D, and viewing results by overlaying the contours.

1.1 Description of the Reaction Model

The reaction that will be simulated is biodegradation of BTEX compounds via different aerobic/anaerobic pathways using multiple electron acceptors. Five different processes considered in this model are Aerobic Respiration (AR), Denitrification (DN), Iron (III) Reduction (IR), Sulfate Reduction (SR), and Methanogenesis (MG). All of these biochemical reactions are assumed to occur in the aqueous phase, mediated by the existing subsurface microbes, and are expected to happen in the following sequence:



The following reaction kinetic framework is used to model the degradation rate of hydrocarbon via different electron acceptor pathways:

$$r_{\text{HC},\text{O}_2} = -k_{\text{HC},\text{O}_2} [\text{HC}] \frac{[\text{O}_2]}{K_{\text{O}_2} + [\text{O}_2]} \dots\dots\dots (2)$$

$$r_{\text{HC},\text{NO}_3} = -k_{\text{HC},\text{NO}_3} [\text{HC}] \frac{[\text{NO}_3]}{K_{\text{NO}_3} + [\text{NO}_3]} \cdot \frac{K_{i,\text{O}_2}}{K_{i,\text{O}_2} + [\text{O}_2]} \dots\dots\dots (3)$$

$$r_{\text{HC},\text{Fe}^{3+}} = -k_{\text{HC},\text{Fe}^{3+}} [\text{HC}] \frac{[\text{Fe}^{3+}]}{K_{\text{Fe}^{3+}} + [\text{Fe}^{3+}]} \cdot \frac{K_{i,\text{O}_2}}{K_{i,\text{O}_2} + [\text{O}_2]} \dots\dots\dots (4)$$

$$\cdot \frac{K_{i,\text{NO}_3}}{K_{i,\text{NO}_3} + [\text{NO}_3]}$$

$$r_{\text{HC},\text{SO}_4} = -k_{\text{HC},\text{SO}_4} [\text{HC}] \frac{[\text{SO}_4]}{K_{\text{SO}_4} + [\text{SO}_4]} \cdot \frac{K_{i,\text{O}_2}}{K_{i,\text{O}_2} + [\text{O}_2]} \dots\dots\dots (5)$$

$$\cdot \frac{K_{i,\text{NO}_3}}{K_{i,\text{NO}_3} + [\text{NO}_3]} \cdot \frac{K_{i,\text{Fe}^{3+}}}{K_{i,\text{Fe}^{3+}} + [\text{Fe}^{3+}]}$$

$$r_{\text{HC},\text{CH}_4} = -k_{\text{HC},\text{CH}_4} [\text{HC}] \frac{[\text{CO}_2]}{K_{\text{CH}_4} + [\text{CO}_2]} \cdot \frac{K_{i,\text{O}_2}}{K_{i,\text{O}_2} + [\text{O}_2]} \dots\dots\dots (6)$$

$$\cdot \frac{K_{i,\text{NO}_3}}{K_{i,\text{NO}_3} + [\text{NO}_3]} \cdot \frac{K_{i,\text{Fe}^{3+}}}{K_{i,\text{Fe}^{3+}} + [\text{Fe}^{3+}]} \cdot \frac{K_{i,\text{SO}_4}}{K_{i,\text{SO}_4} + [\text{SO}_4]}$$

where r_{HC,O_2} is the rate at which hydrocarbon is destroyed by utilizing oxygen, $r_{\text{HC},\text{NO}_3}$ is the rate at which hydrocarbon is destroyed by utilizing Nitrate, $r_{\text{HC},\text{Fe}^{2+}}$ is the rate at which hydrocarbon is destroyed by producing Fe^{2+} (or utilizing Fe^{3+}), $r_{\text{HC},\text{SO}_4}$ is the rate at which hydrocarbon is destroyed by utilizing sulfate, $r_{\text{HC},\text{CH}_4}$ is the rate at which hydrocarbon is destroyed by producing methane, $[\text{O}_2]$ is the oxygen concentration $[\text{ML}^{-3}]$, k_{O_2} is the first-order rate constant $[\text{T}^{-1}]$, K_{O_2} is the Monod half saturation constant $[\text{ML}^{-3}]$ (by setting all the half-saturation constants to a small value, zero-order dependency can be simulated with respect to the electron donor and hence a first-order degradation model with respect to hydrocarbon; the default values simulate this option), and K_{i,O_2} is the oxygen inhibition constant $[\text{ML}^{-3}]$ (by setting all inhibition constants to a small value, reactions can be forced to occur in a sequential fashion; the default values simulate this process). Similar nomenclature is used to identify other Monod constants and inhibition coefficients.

Since the concentrations of Fe^{3+} and CO_2 are not readily measurable under normal field conditions, these terms were replaced with the “assimilative capacity” for iron reduction and methanogenesis, defined as the following:

$$[\text{Fe}^{3+}] = [\text{Fe}^{2+}_{\text{max}}] - [\text{Fe}^{2+}] \dots\dots\dots (7)$$

$$[MC]=[CO_2] = [CH_{4,max}] - [CH_4] \dots\dots\dots (8)$$

where $[Fe^{2+}_{max}]$ and $[CH_{4,max}]$ are the maximum possible aquifer levels of these species that represent the aquifer's maximum capacity for iron reduction and methanogenesis. Note: the concentration of CO_2 used here is the CO_2 evolved while the hydrocarbon is destroyed via methanogenesis, which may be thought of as the "Methanogenic Capacity" (MC) of the aquifer. Using these relations, iron (III) reduction and methanogenesis processes may be related back to measurable Fe^{2+} and CH_4 concentration levels.

The total rate of hydrocarbon destruction, via all the above described processes, is the sum of each of the individual rates and is given as the following:

$$\frac{d[HC]}{dt} = r_{HC,O_2} + r_{HC,NO_3} + r_{HC,Fe^{2+}} + r_{HC,SO_4} + r_{HC,CH_4} \dots\dots\dots (9)$$

Rates of electron acceptor utilization are given as the corresponding rate of hydrocarbon destruction multiplied by the appropriate yield coefficient (Y):

$$\frac{d[O_2]}{dt} = Y_{O_2/HC} r_{HC,O_2} \dots\dots\dots (10)$$

$$\frac{d[NO_3]}{dt} = Y_{NO_3/HC} r_{HC,NO_3} \dots\dots\dots (11)$$

$$\frac{d[Fe^{2+}]}{dt} = -Y_{Fe^{2+}/HC} r_{HC,Fe^{2+}} \dots\dots\dots (12)$$

$$\frac{d[SO_4]}{dt} = Y_{SO_4/HC} r_{HC,SO_4} \dots\dots\dots (13)$$

$$\frac{d[CH_4]}{dt} = -Y_{CH_4/HC} r_{HC,CH_4} \dots\dots\dots (14)$$

The yield values (the mass ratio of electron acceptors removed or metabolic byproducts produced to total BTEX degraded) are as follows: $Y_{O_2/HC} = 3.14$, $Y_{NO_3/HC} = 4.9$, $Y_{Fe^{2+}/HC} = 21.8$, $Y_{SO_4/HC} = 4.7$, and $Y_{CH_4/HC} = 0.78$. Typical values of all inhibition coefficients, except for $K_{iFe^{3+}}$, should be in the range of 1.0 to 0.01 mg/L. $K_{iFe^{3+}}$ should always be set around 40% to 80% of the max Fe^{2+} value. Monod half-saturation constants should be in the range of 1.0 to 0.1 mg/L.

As pointed out in Lu et al.,¹ it is important to note that this model is based on several assumptions. The model should be used with caution only at sites where these assumptions are valid.

In summary, the key assumptions used in the model are: (1) the fuel chemical species benzene, toluene, ethylbenzene, and xylene are assumed to degrade at similar rates and, hence, are combined and modeled as a single electron donor species BTEX; (2) production of Fe^{2+} and methane are restricted at a node to a “maximum-observed level”; however, the model assumes that an infinite supply of electron acceptors will be available for iron-reduction and methanogenic reactions; (3) more complex processes are not considered; these include processes such as the rate-limited interaction of bioavailable, solid-phase Fe^{3+} and aqueous-phase Fe^{2+} , interaction of oxygen and Fe^{2+} , and/or variations in the spatial pattern of methanogenic activity and CO_2 availability; (4) growth and decay of various microbial populations and their interactions with contaminants and aquifer solids are assumed to be negligible; and (5) all BTEX decay reactions are approximated as first-order reactions, and, hence, the model ignores the Monod limitation due to the electron donor (BTEX) availability. Fortunately, these assumptions are expected to be reasonable approximations for most field sites. However, there will always be some exceptions.

1.2 Description of Problem

The problem in this tutorial is the same as the problem described in “RT3D – Instantaneous Aerobic Degradation” tutorial. The site is a 510 m x 310 m section of a confined aquifer with a flow gradient from left to right. An underground storage tank is leaking fuel hydrocarbon contaminants at 2 m³/day at the location shown. Concentration of BTEX is 1000 mg/L. Initial levels of hydrocarbon, oxygen, nitrate, Fe^{2+} , sulfate, and methane in the aquifer are assumed to be as follows: 0.0, 4.0, 20.0, 0.0, 10.0, and 0.0 mg/L, respectively. The reaction constants to be used are as follows:

Constant	Value
Max Fe^{2+}	25.0 mg/L
Max Methane	30.0 mg/L
k_{HC,O_2}	0.1 day ⁻¹
$k_{\text{HC},\text{NO}_3}$	0.008 day ⁻¹
$k_{\text{HC},\text{Fe}^{3+}}$	0.0005 day ⁻¹
$k_{\text{HC},\text{SO}_4}$	0.00025 day ⁻¹
$k_{\text{HC},\text{CH}_4}$	0.0001 day ⁻¹

Values of all inhibition coefficients, except $K_{\text{ifFe}^{3+}}$, are assumed to be 0.01 mg/L and all Monod constants are assumed to 0.5 mg/L. The value of $K_{\text{ifFe}^{3+}}$ is set at 12.5 mg/L (50% of max Fe). The yield values are the same as those described in the previous section.

¹ Lu, G., Clement, T.P., Zheng, C., and Wiedemeier, T.H. (1999). Natural attenuation of BTEX compounds: Model development and field-scale application. *Ground Water*. 37(5):707–717.

The first part of the problem will be to import a previously computed MODFLOW flow model of the site. Using this flow field, a reactive transport model will then be defined using RT3D.

1.3 Getting Started

Do the following to get started:

1. If GMS is not running, launch GMS.
2. If GMS is already running, select the *File* / **New** command to ensure the program settings are restored to the default state.

2 Importing the MODFLOW Model

The first part of the simulation is to import the MODFLOW flow model. A steady state flow model has been previously computed and is supplied with the tutorial files.

1. Select the *File* | **Open** command to bring up the *Open* dialog.
2. Locate the directory entitled *Tutorials\RT3D\flowmod*.
3. Select the file entitled “flowmod.gpr”.
4. Click **Open** to import the project file.

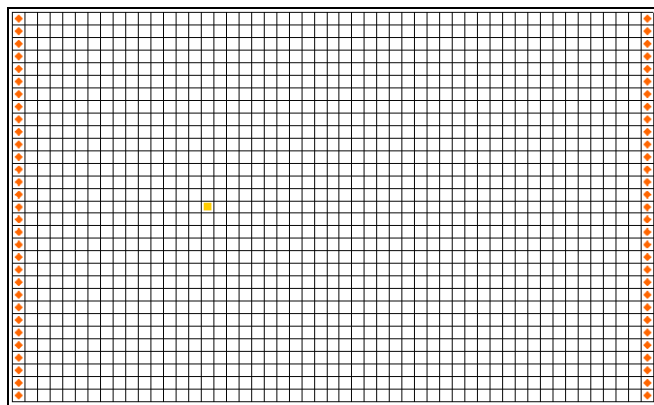




Figure 1 The MODFLOW flow model

3 Building the Transport Model

Now that the flow model is imported, the next step is to set up the RT3D simulation. For this part of the simulation, select the reaction, define the reaction data, define the supplemental layer data needed by RT3D, and assign concentrations to the well.

4 Initializing the Model

The first step is to initialize the RT3D data.

1. Expand the “ 3D Grid Data” folder in the Project Explorer.
2. Right-click on the “ grid” item in the Project Explorer.
3. Select the **New MT3DMS...** command to open the *Basic Transport Package* dialog.

5 The BTN Package

The next step is to initialize the data in the Basic Transport Package. First, initialize the data, select RT3D as the transport model, and select the appropriate packages.

1. In the *Model* section, select the *RT3D* option.
2. Select the **Packages** button to open the *MT3DMS/RT3D Packages* dialog.
3. Select the following packages:
 - *Advection Package*
 - *Dispersion Package*
 - *Source/Sink Mixing Package*
 - *Chemical Reaction Package*
4. For the RT3D reactions, select the “Kinetic-Limited Deg. of BTEX w/ MEA” option.
5. Select the **OK** button to exit the *MT3DMS/RT3D Packages* dialog.

5.1 Starting Concentrations

Note that in the *Layer Data* section of the dialog, the species associated with the reaction being modeled are listed by name. The next step is to define the starting concentration for each of these species. The default starting concentration is zero. Use the default value for BTEX, Fe²⁺, and methane, but use non-zero values for oxygen (4.0 mg/L), nitrate (20.0 mg/L), and sulfate (10.0 mg/L).

1. Select *Oxygen* from the list of species in the spread sheet.
2. Enter a value of “4.0” for the starting concentration.
3. Enter a starting value of “20.0” for *Nitrate*.

4. Enter a starting value of “10.0” for *Sulfate*.

5.2 Porosity

The porosity to be used is 0.3. Since this is the default in GMS, nothing needs to be changed.

5.3 Stress Periods

Next, define the stress periods. Since the injection rate of the well and the other boundary conditions do not change, use a single stress period with a length of 730 days (two years).

1. Select the **Stress Periods** button to open the *Stress Periods* dialog.
2. Enter a value of “730” for the *Length*.
3. Enter a value of “10” for the *Number of time steps*.
4. Select the **OK** button to exit the *Stress Periods* dialog.

5.4 Output Options

Finally, define the output options. One binary solution file is created by RT3D for each of the species. By default, RT3D saves a solution at each transport step for each species. Since this results in large files containing more solutions than needed for the simple post-processing, specify that a solution be saved every 73 days (every time step).


1. Select the **Output Control...** button to open the *Output Control* dialog.
2. Select the *Print or save at specified times* option.
3. Select the **Times** button to open the *Variable Time Steps* dialog.
4. Select the **Initialize Values** button to open the *Initialize Time Steps* dialog.
5. Enter “73.0” for the *Initial time step size*.
6. Enter “73.0” for the *Maximum time step size*.
7. Enter “730.0” for the *Maximum simulation time*.
8. Select the **OK** button to exit the *Initialize Time Steps* dialog.
9. Select the **OK** button to exit the *Variable Time Steps* dialog.
10. Select the **OK** button to exit the *Output Control* dialog.

This completes the input for the Basic Transport package.

11. Select the **OK** button to exit the *Basic Transport Package* dialog.

6 Assigning Concentrations to the Left Boundary

The left boundary of the model is a constant head boundary. Since the head at the left boundary is greater than the head at the right boundary, the left boundary acts as a source and water enters the model from the left. Thus, the concentrations of the species at the left boundary must be defined. The simplest way to do this is to mark the cells as specified concentration cells.

1. Using the **Select Cell**  tool, select the column of cells on the left boundary by dragging a box that just surrounds the cells.
2. Right-click on any of the selected cells and select the **Properties...** command from the pop-up menu to open the *3D Grid Cell Properties* dialog.
3. Switch to the *RT3D* tab.
4. Change the *ICBUND* value to “-1”.
5. Select the **OK** button to close the *3D Grid Cell Properties* dialog.

7 The Advection Package

The next step is to initialize the data for the Advection package.

1. Select the *RT3D* | **Advection Package** command to open the *Advection Package* dialog.
2. Select the “Standard finite-difference method” solver from the *Solution scheme* drop-down menu of the dialog.
3. Select the **OK** button to close the *Advection Package* dialog.

8 The Dispersion Package


Next, enter the data for the Dispersion package. The aquifer has a longitudinal dispersivity of 10.0 m and a transverse (horizontal) dispersivity of 3.0 m. The vertical dispersivity is assumed to be equal to the longitudinal dispersivity.

1. Select the *RT3D* | **Dispersion Package** command to open the *Dispersion Package* dialog.
2. Select the **Longitudinal Dispersivity** button to open the *Longitudinal Dispersivity* dialog.

3. Select the **Constant** → **Grid** button to open the *Grid Value* dialog.
4. Enter a value of “10.0”.
5. Select **OK** to close the *Grid Value* dialog.
6. Select the **OK** button to exit the *Longitudinal Dispersivity* dialog.
7. Enter a value of “0.3” for the *TRPT* value.
8. Select the **OK** button to exit the *Dispersion Package* dialog.

9 The Source/Sink Mixing Package

Next, define the concentration at the spill location. Assign a concentration of 1000 mg/L for BTEX and leave the other electron acceptor concentrations at the default value of zero.

1. Using the **Select Cell**  tool, select the cell containing the injection well (spill location) by clicking anywhere in the interior of the cell.
2. Select the *RT3D* | **Point Sources/Sinks** command to open the *MODFLOW/RT3D Source/Sinks* dialog.
3. Click the **Add BC** button near the bottom of the dialog.
4. Change the *Type (ITYPE)* to “well (WEL)”.
5. In the *BTEX* column, enter a concentration of “1000” (mg/L).
6. Select the **OK** button to exit the *MODFLOW/RT3D Source/Sinks* dialog.

10 The Chemical Reaction Package

Next, initialize the Chemical Reaction package and define appropriate reaction rate constants.

1. Select the *RT3D* | **Chemical Reaction Package** command to open the *RT3D Chemical Reaction Package* dialog.
2. In the *Reaction parameters* section, click on the *Max_Fe2+* item and set its value to “25.0”.
3. Set the value of *Max Methane* to “30.0”.
4. Set the value of k_{HC,O_2} to “0.1”.
5. Set the value of k_{HC,NO_3} to “0.008”.

6. Set the value of $k_{HC,Fe^{3+}}$ to “0.0005”.
7. Set the value of k_{HC,SO_4} to “0.00025”.
8. Set the value of k_{HC,CH_4} to “0.0001”.

All monod constants inhibition coefficients, and yield values, except $K_{iFe^{3+}}$, will be left at the default values. The value of $K_{iFe^{3+}}$ is set at 12.5 mg/L (50% of max Fe).

9. Set the value of $K_{iFe^{3+}}$ to “12.5”.
10. Select the **OK** button to exit the *RT3D Chemical Reaction Package* dialog.

11 Running RT3D

At this point, it is possible to save the model and run RT3D.

1. Select the *File* | **Save As** command to bring up the *Save As* dialog.
2. Locate and open the directory entitled *Tutorials\RT3D\btexkin*.
3. Enter “btmod” for the file name.
4. Select the **Save** button.

Since the name of the project has changed, either use the **Run Options** command to tell RT3D where the MODFLOW solution is, or re-run MODFLOW. In this case, do the latter.

5. Select the *MODFLOW* | **Run MODFLOW** command to start the *MODFLOW* model wrapper.
6. When the simulation is finished, select the **Close** button to exit the *MODFLOW* model wrapper. The solution will automatically be imported.


To run RT3D:

7. Select the *RT3D* | **Run RT3D** command to start the *RT3D* model wrapper.
8. Select **Yes** at the prompt to save the changes.
9. When the simulation is finished, select the **Close** button to exit the *RT3D* model wrapper.

GMS reads in the solution automatically.

12 Viewing the Results

First, view the BTEX solution at 730 days.

1. Select the “ BTEX” dataset from the Project Explorer. (if needed, expand the “btmod (RT3D)” solution in the Project Explorer.)
2. Select the time step at 730 days from the *Time Steps Window*.

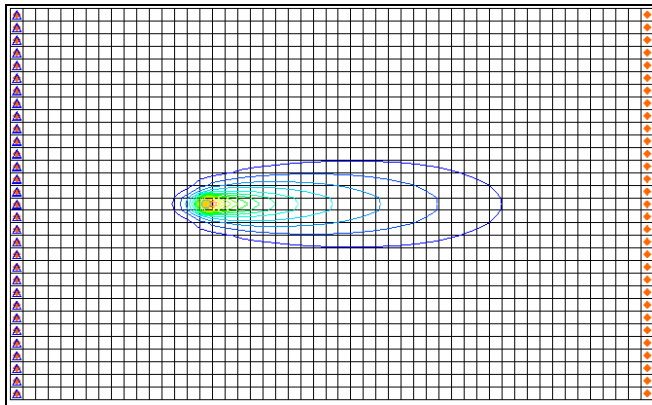



Figure 2 BTEX solution at 730 days

12.1 Overlaying Contours

Next, view concentration contours for the other species. When working with multiple species, it is often convenient to overlay contours of one species on top of contours from another species. As an example, overlay the contours of the other species on top of the BTEX contours. To do this, display the BTEX contours in the background as a CAD layer.



Before creating the CAD layer, turn off most of the display options other than the contours (to avoid saving the grid cell lines to the CAD file).

1. Select the **Display Options**  macro to open the *Display Options* dialog.
2. Turn off the *Cells edges* option.
3. Switch to the *RT3D* tab
4. Select the **Check None** button.
5. Switch to the *MODFLOW* tab.
6. Select the **Check None** button.
7. Select **OK** to close the *Display Options* dialog.


To create the CAD layer:

8. In the Project Explorer, right-click on the empty space then, from the pop-up menu, select the **Convert to CAD** command to open the *Save CAD data* dialog.
9. Enter “BTEX_contours” as the file name.
10. Select **Save** to create the CAD data.

Now reset the display options and view the oxygen solution at 730 days.

11. Select the **Display Options**  macro to open the *Display Options* dialog.
12. Turn on the *Layer Borders* option.
13. Select the **OK** button to close the *Display Options* dialog.
14. Select the “ Oxygen” dataset from the Project Explorer.

Now turn on a color ramp and a legend.

15. Select the **Contour Options**  macro to open the *Dataset Contour Options – 3D Grid – Oxygen* dialog.
16. Select the “Color Fill” option for the *Contour Method*.
17. Change the *Transparency* value to “50”.
18. Select the **OK** button to exit the *Dataset Contour Options – 3D Grid – Oxygen* dialog.
19. Select the time step at 730 days from the *Time Steps Window*.

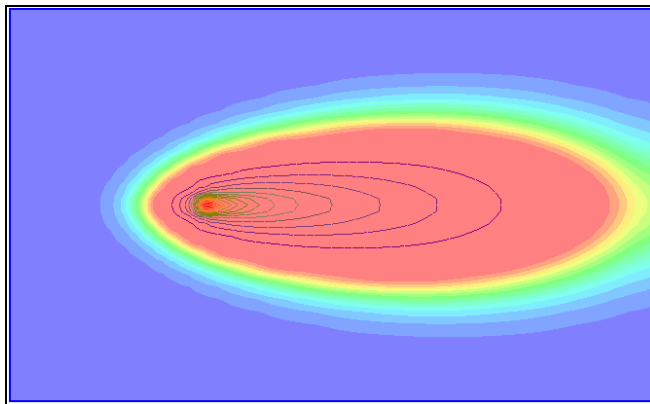




Figure 3 The Oxygen solution at 730 days

Next, view the nitrate solution at 730 days.

20. Select the “ Nitrate” dataset from the Project Explorer.
21. Select the time step at 730 days from the *Time Steps Window*.

Finally, view the Fe^{2+} solution at 730 days.

22. Select the “ Iron (Fe^{2+})” dataset from the Project Explorer.
23. Select the time step at 730 days from the *Time Steps Window*.

13 Other Post-Processing Options

At this point, feel free to experiment with the other post-processing options, including film loop animation.

14 Conclusion

This concludes the tutorial. Continue to explore RT3D applications or close the program.